

Bentonite Functionalised with 2-(3-(2-aminoethylthio)propylthio)ethanamine (AEPE) for the Removal of Hg(II) from Wastewaters



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Reviews

It is great and fantastic. Better then never, though i am quite late in start reading this one. Its been written in an extremely simple way and is particularly only right after i finished reading this ebook where actually changed me, affect the way i really believe.

(Orin Blick)

BENTONITE FUNCTIONALISED WITH 2-(3-(2-AMINOETHYLTHIO)PROPYLTHIO)ETHANAMINE (AEPE) FOR THE REMOVAL OF Hg(II) FROM WASTEWATERS



Diplomica Verlag Nov 2008, 2008. Taschenbuch. Book Condition: Neu. 270x190x5 mm. This item is printed on demand - Print on Demand Neuware - In this study, natural bentonite clay was first purified and then functionalised with the chelating ligand 2-(3-(2-aminoethylthio)propylthio)ethanamine (AEPE) to improve the adsorption capacity and selectivity towards Hg(II) ions. The surface modification was characterised with the help of powder X-ray diffraction (XRD), Fourier transform infrared spectroscopy (FT-IR), BET isotherm to determine the specific surface area while the thermal stability of the samples was studied using thermogravimetric analysis (TGA). FT-IR and TGA demonstrated the presence of the chelating ligand on the modified clay. XRD pattern indicated that the chelating agent AEPE was only grafted onto the surface of the purified bentonite, whereas the interlayer distance did not change. N₂ specific surface area measurement also indicated the coverage of AEPE onto the surface of purified bentonite. Adsorption of Hg(II) ions from aqueous solutions as a function of pH, contact time, initial concentration, ionic strength, interfering ions and adsorbent dose was studied. The adsorption process followed a pseudo-second order kinetics and monolayer adsorption. The adsorption of Hg(II) ions increased with increasing pH and reached a plateau value in the pH range of 4.0-8.0. The removal of Hg(II) was found to be higher than 99% at an initial concentration of 20 mg/L using adsorbent dose of 0.01 g. The presence of NaNO₃ as background electrolytes at concentration ranging from 0.01 to 2.0 M decreased the adsorption of Hg(II) ions. Furthermore, the adsorption capacity increased with increasing adsorbent dose. Sorption data analysis was carried out using Langmuir isotherm for the uptake of Hg(II) ions at a concentration range of 20-400 mg/L. The adsorption process was found to be favourable as the separation parameter is less than unity (RL1). The maximal adsorption capacity was...



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